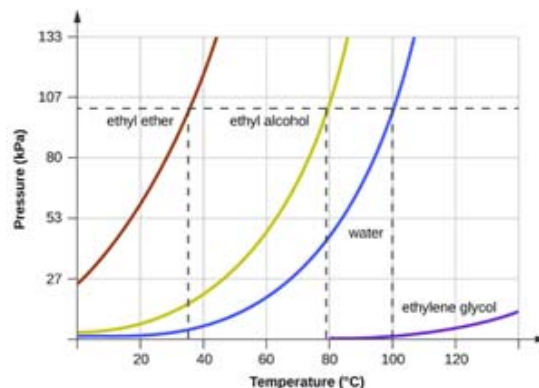


## Assignment 7: Phase Change/Temperature Change Worksheet

1. Use table to the right to find the boiling point of water at 80 kPa.



2. Find the heat needed to raise 250g of ice @ -15 °C to liquid water @ 25 °C.
3. Find the heat lost when 45.5 g of water is cooled from 458 K to 333 K @ 101.3 kPa.
4. How much water can be vaporized by 230 kJ?
5. How much liquid water is produced from a 2.6 kg block of ice that is originally @ -12 °C with 65 kcal of heat?
6. A tub is fill with 200 kg of water @ 14.8 °C. A 100 kg block of hot aluminum is dropped into the tub. Assume no water evaporated and the final temperature of water/ion combination is 22 °C. How hot was the aluminum? Aluminum has a specific heat of 0.22 cal/g C°.