Chemistry: Unit 5 Assignment 2

Name:

1. Balance the following equations

a.
$$Al_{(s)} + FeO_{(s)} \rightarrow Al_2O_{3(s)} + Fe_{(l)}$$

b. ___ FeCl_{3(aq)} + ___ NH₄OH_(aq)
$$\Rightarrow$$
 ___ Fe(OH)_{3(s)} + ___ NH₄Cl_(aq)

c.
$$\underline{\hspace{1cm}}$$
 Sn_(s) + $\underline{\hspace{1cm}}$ HNO_{3(aq)} \rightarrow SnO_{2(s)} + $\underline{\hspace{1cm}}$ NO_{2(g)} + $\underline{\hspace{1cm}}$ H₂O_(l)

- 2. Write the molecular, ionic and net ionic equations for the following reactions. You will need to decide the state of the products.
 - a. Aqueous sodium phosphate reacts with aqueous silver nitrate to form sodium nitrate and silver phosphate

b. Magnesium metal reacts with an aqueous cuprous nitrate to produce magnesium nitrate and copper metal

c. Lithium metal reacts with oxygen to produce lithium oxide