

A. Define the following terms

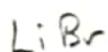
1. Oxidation state - charge of a stable ion
2. Cation - positive ion; formed by losing e^-
3. Anion - negative ion; formed by gaining e^-
4. Ionic Bond - electrostatic attraction between oppositely charged ions
5. Electronegativity - pull on bonded electrons

B. Identify the oxidative state of the following atoms/ions in their stable state (write the symbol and charge for each)

6. Magnesium	7. Gallium	8. Sulfur	9. Bromine	10. Argon
Mg^{2+}	Ga^{3+}	S^{2-}	Br^{-}	Ar^0
11. Iron (III)	12. Cupric	13. Aurous	14. Nitrogen	15. Potassium
Fe^{3+}	Cu^{2+}	Au^{1+}	N^{3-}	K^{1+}

Write the correct formula and name the compounds for the following sets of atoms/ions

16. Lithium and Bromine



lithium bromide

19. Iron (III) and Fluorine



iron(III) fluoride

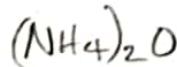
ferric fluoride

22. Strontium and Nitrite



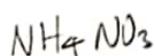
strontian nitrate

25. Ammonium and Oxygen



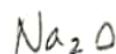
ammonium oxide

28. Nitrate and Ammonium



ammonium nitrate

17. Sodium and Oxygen



sodium oxide

20. Cuprous and hydroxide



copper(I) hydroxide

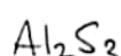
cuprous hydroxide

23. Zinc and Perchlorate



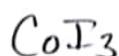
zinc perchlorate

26. Aluminum and Sulfur



aluminum sulfide

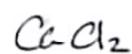
29. Cobalt(III) and Iodine



cobalt(III) iodide

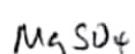
cobaltic iodide

18. Chlorine and Calcium



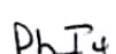
calcium chloride

21. Sulfate and Magnesium



magnesium sulfate

24. Iodine and Lead (IV)



lead(IV) iodide

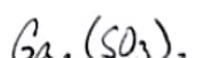
plumbic iodide

27. Barium and Carbonate



barium carbonate

30. Sulfite and Gallium



gallium sulfite