

1. Name the following compounds

a. CaBr_2

calcium bromide

b. Au_2O_3 gold (III) oxide
auric oxidec. N_2O

dinitrogen monoxide

d. $\text{HIO}_3(\text{aq})$

iodic acid

2. Write the formulas for compounds from the following elements/ions

a. aluminum & nitrogen



b. zinc & oxide



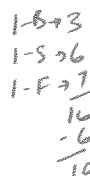
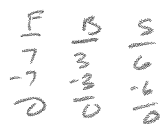
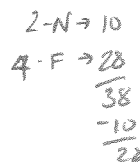
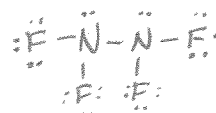
c. oxygen & fluorine



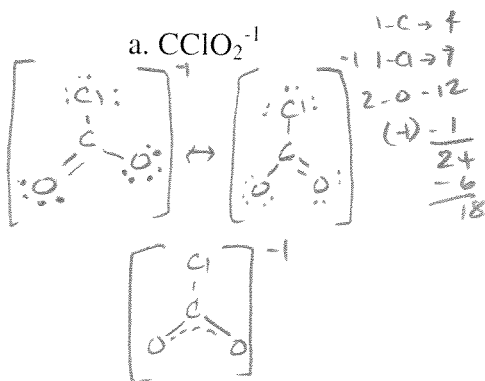
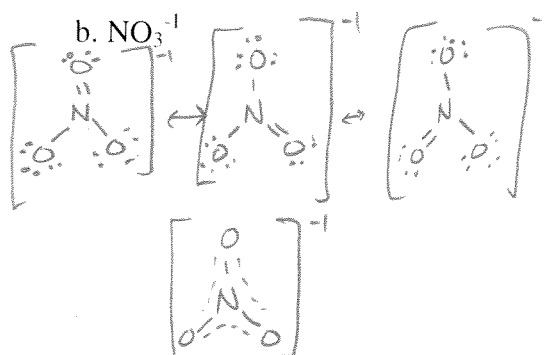
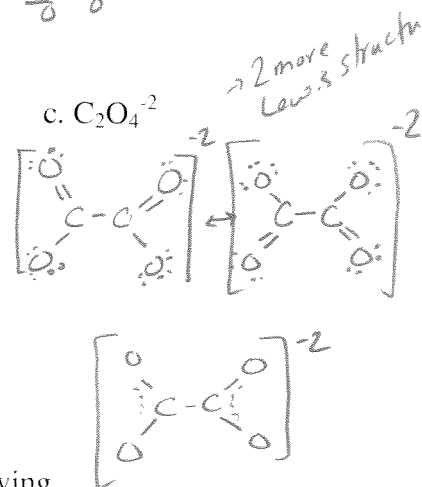
d. sulfite & gallium



3. Draw Lewis structures and calculate formal charge for each atom in the structure.

a. HI b. CO c. BSF d. N_2F_4 

4. For the following, draw the Lewis structures and the resonance structure

a. CClO_2^{-1} b. NO_3^{-1} c. $\text{C}_2\text{O}_4^{-2}$ 

5. Draw the Lewis structure, write the VSEPR and geometric structure for the following

a. H_2S b. PCl_5 